

# Cambridge IGCSE™

CANDIDATE  
NAME

CENTRE  
NUMBER

--	--	--	--	--

CANDIDATE  
NUMBER

--	--	--	--

## CHEMISTRY

0620/31

Paper 3 Theory (Core)

May/June 2023

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **16** pages.





- 2 (a) Table 2.1 shows some properties of the halogens.

**Table 2.1**

halogen	melting point in °C	boiling point in °C	density at room temperature and pressure in g/cm <sup>3</sup>
fluorine	-220	-188	0.0016
chlorine	-101	-35	0.0032
bromine		+59	3.1
iodine	+114	+184	

Use the information in Table 2.1 to predict:

- (i) the melting point of bromine

..... [1]

- (ii) the density of iodine at room temperature and pressure

..... [1]

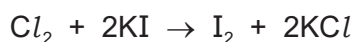
- (iii) the physical state of chlorine at -10 °C. Give a reason for your answer.

physical state .....

reason .....

..... [2]

- (b) The equation for the reaction of aqueous chlorine with aqueous potassium iodide is shown.



- (i) Choose the word which best describes this type of chemical reaction.  
Draw a circle around your chosen answer.

**addition**      **displacement**      **neutralisation**      **polymerisation**      [1]

- (ii) Explain why aqueous iodine does **not** react with aqueous potassium chloride.

..... [1]

(c) Complete the diagram in Fig. 2.1 to show the electronic configuration of a chlorine atom.

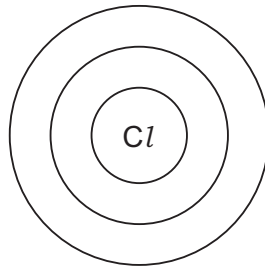


Fig. 2.1

[1]

(d) Describe a test for chlorine.

test .....

observations .....

[2]

[Total: 9]

- 3 (a) Water from natural sources contains dissolved gases.

Choose from the list, the gas that is essential for aquatic life.  
Draw a circle around your chosen answer.

argon      hydrogen      nitrogen      oxygen      [1]

- (b) Polluted water may contain harmful substances such as metal compounds, plastics, nitrates and phosphates.

(i) Name one **other** harmful substance which is present in polluted water.

..... [1]

(ii) State why nitrates are harmful to aquatic life.

..... [1]

- (c) Table 3.1 shows the masses of ions, in mg, present in a 1000 cm<sup>3</sup> sample of polluted water.

**Table 3.1**

name of ion	formula of ion	mass of ion present in mg/ 1000 cm <sup>3</sup> of polluted water
	NH <sub>4</sub> <sup>+</sup>	0.5
calcium	Ca <sup>2+</sup>	2.2
chloride	Cl <sup>-</sup>	2.5
hydrogencarbonate	HCO <sub>3</sub> <sup>-</sup>	12.0
magnesium	Mg <sup>2+</sup>	0.8
nitrate	NO <sub>3</sub> <sup>-</sup>	0.4
potassium	K <sup>+</sup>	8.3
silicate	SiO <sub>3</sub> <sup>2-</sup>	8.0
sodium	Na <sup>+</sup>	10.2
sulfate	SO <sub>4</sub> <sup>2-</sup>	0.2
tin(II)	Sn <sup>2+</sup>	0.4

Answer these questions using information from Table 3.1.

(i) Name the negative ion present in the highest concentration.

..... [1]

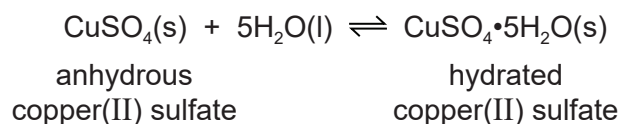
(ii) State the name of the NH<sub>4</sub><sup>+</sup> ion.

..... [1]

(iii) Calculate the mass of calcium ions present in 200 cm<sup>3</sup> of polluted water.

mass = ..... mg [1]

(d) Copper(II) sulfate can be used to test for the presence of water.



(i) State the meaning of the term hydrated.

..... [1]

(ii) Describe how hydrated copper(II) sulfate is changed to anhydrous copper(II) sulfate.

..... [1]

(e) Complete the symbol equation for the reaction of sodium with water.



[Total: 10]

4 This question is about sulfur and compounds of sulfur.

(a) Sulfur has several isotopes.

Define the term isotopes.

.....  
 ..... [2]

(b) Deduce the number of protons, neutrons and electrons in the sulfide ion shown.



number of protons .....

number of neutrons .....

number of electrons .....

[3]

(c) Sulfur burns in oxygen to produce sulfur dioxide.

Fig. 4.1 shows an incomplete reaction pathway diagram for this reaction.

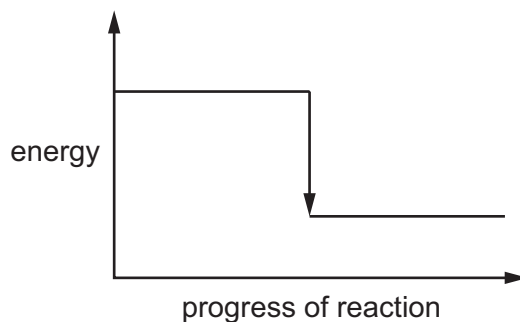


Fig. 4.1

(i) Complete Fig. 4.1 by writing these formulae on the diagram:

- S + O<sub>2</sub>
- SO<sub>2</sub>.

[1]

(ii) Explain how Fig. 4.1 shows that the reaction is exothermic.

.....  
 ..... [1]

(iii) Complete this sentence about an exothermic reaction using a word from the list.

**products      reactants      sulfur      surroundings**

An exothermic reaction transfers thermal energy to the ..... [1]

- (d) Fig. 4.2 shows the apparatus used for the electrolysis of dilute sulfuric acid using graphite electrodes.

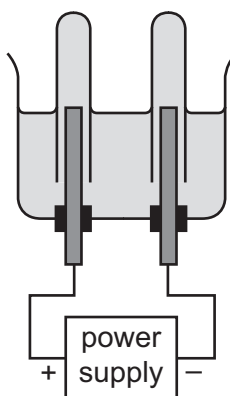


Fig. 4.2

- (i) Label Fig. 4.2 to show:
- the cathode
  - the electrolyte.
- [2]

- (ii) Name the products and state the observations at the positive and negative electrodes.

product at the positive electrode

.....

observations at the positive electrode

.....

product at the negative electrode

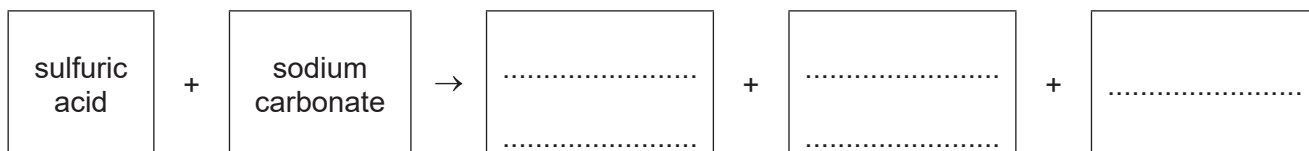
.....

observations at the negative electrode

.....

[4]

- (e) Complete the word equation for the reaction of dilute sulfuric acid with sodium carbonate.



[3]

- (f) A few drops of thymolphthalein indicator are added to dilute sulfuric acid.

State the colour of the solution.

..... [1]

[Total: 18]

5 This question is about metals.

(a) Iron is a transition element. Potassium is an element in Group I of the Periodic Table.

State **two** differences in the physical properties of iron compared to potassium.

1 .....

2 ..... [2]

(b) Carbon is used to extract iron from iron ore in a blast furnace.

State **two** uses of carbon in the extraction process.

1 .....

2 ..... [2]

(c) Steel is an alloy of iron.

(i) State the meaning of the term alloy.

.....

..... [1]

(ii) State why alloys are more useful than pure metals.

..... [1]

(d) Table 5.1 shows the observations made when four different metals react with dilute hydrochloric acid of the same concentration.

**Table 5.1**

metal	observations
iron	bubbles form slowly
lead	no bubbles formed
magnesium	bubbles form rapidly
nickel	bubbles form very slowly

Put the four metals in order of their reactivity.

Put the least reactive metal first.

least reactive  $\longrightarrow$  most reactive

--	--	--	--

[2]

[Total: 8]

- 6 (a) A student investigates the reaction of small pieces of zinc of the same mass and size with three different concentrations of dilute hydrochloric acid in the presence of a catalyst.

The three concentrations of dilute hydrochloric acid are:

- 1.0 mol/dm<sup>3</sup>
- 1.5 mol/dm<sup>3</sup>
- 2.0 mol/dm<sup>3</sup>.

All other conditions stay the same.

Table 6.1 shows the time taken for each reaction to finish.

**Table 6.1**

concentration of hydrochloric acid in mol/dm <sup>3</sup>	time taken for the reaction to finish in s
	200
	100
	150

- (i) Complete Table 6.1 by writing the concentrations of hydrochloric acid in the first column. [1]

- (ii) Describe the effect on the time taken for the zinc to finish reacting with 2.0 mol/dm<sup>3</sup> hydrochloric acid with no catalyst present.

All other conditions stay the same.

..... [1]

- (iii) Describe the effect on the time taken for the zinc to finish reacting with 2.0 mol/dm<sup>3</sup> hydrochloric acid when the surface area of the zinc is increased.

All other conditions stay the same.

..... [1]

- (b) Crystals of zinc chloride can be prepared by reacting excess zinc with dilute hydrochloric acid.

Choose from the list, the method used to separate the unreacted zinc from the reaction mixture.

Draw a circle around your chosen answer.

**chromatography      crystallisation      evaporation      filtration      [1]**

(c) Zinc chloride is soluble in water.

Choose one **other** compound that is soluble in water.

Tick (✓) **one** box.

calcium carbonate

lead(II) chloride

silver chloride

sodium nitrate

[1]

[Total: 5]

7 (a) Fig. 7.1 shows the displayed formula of mesaconic acid.

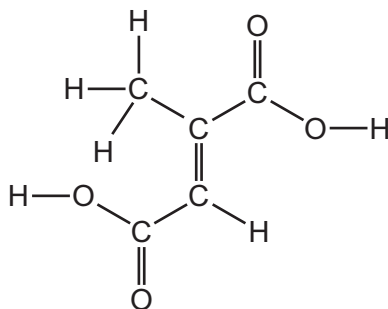


Fig. 7.1

(i) On Fig. 7.1 draw a circle around **one** carboxylic acid functional group. [1]

(ii) Deduce the molecular formula of mesaconic acid.

..... [1]

(iii) Mesaconic acid is a colourless compound.

Describe the colour change when excess mesaconic acid is added to aqueous bromine.

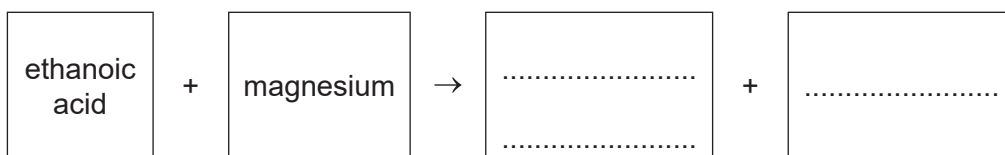
from ..... to ..... [2]

(b) Ethanoic acid belongs to the homologous series of carboxylic acids.

Define the term homologous series.

.....  
 ..... [2]

(c) Complete the word equation for the reaction of ethanoic acid with magnesium.



[2]

- (d) Ethanoic acid reacts with ethanol.  
The organic product has the molecular formula  $C_4H_8O_2$ .

Complete Table 7.1 to calculate the relative molecular mass of  $C_4H_8O_2$ .

**Table 7.1**

atom	number of atoms	relative atomic mass	
carbon	4	12	$4 \times 12 = 48$
hydrogen		1	
oxygen		16	

relative molecular mass = ..... [2]

- (e) Ethanol can be manufactured by fermentation.

Complete the word equation for one **other** method of manufacturing ethanol.

..... + .....  $\rightarrow$  ethanol [2]

[Total: 12]

8 This question is about nitrogen and compounds of nitrogen.

(a) Nitrogen is a non-metal. Nitrogen is a poor electrical conductor.

Describe two **other** physical properties which are typical of non-metals.

1 .....

2 .....

[2]

(b) Oxides of nitrogen are air pollutants which contribute to acid rain.

(i) State **one** source of oxides of nitrogen in the air.

..... [1]

(ii) State one **other** adverse effect of oxides of nitrogen.

..... [1]

(c) Ammonia is a simple molecule with covalent bonds.

(i) Describe a covalent bond.

.....

..... [2]

(ii) Complete Fig. 8.1 to show the dot-and-cross diagram for a molecule of ammonia.

Show outer shell electrons only.

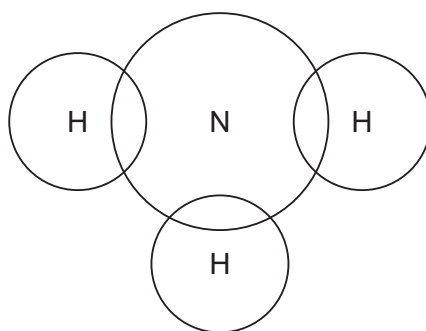


Fig. 8.1

[2]

(iii) Aqueous ammonia is alkaline.

Choose from the list, the pH value that is alkaline.

Draw a circle around your chosen answer.

pH 1

pH 5

pH 7

pH 10

[1]

(iv) Aqueous ammonia releases ammonia gas.

Ammonia gas turns damp red litmus paper blue.

A long glass tube is set up as shown in Fig. 8.2.

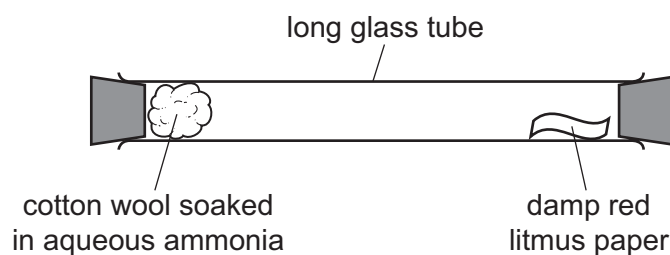


Fig. 8.2

At first, the litmus paper does **not** turn blue.

After a short time, the litmus paper turns blue.

Explain these results in terms of the kinetic particle theory.

.....

.....

.....

.....

.....

..... [3]

[Total: 12]

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge Assessment International Education Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cambridgeinternational.org](http://www.cambridgeinternational.org) after the live examination series.

Cambridge Assessment International Education is part of Cambridge Assessment. Cambridge Assessment is the brand name of the University of Cambridge Local Examinations Syndicate (UCLES), which is a department of the University of Cambridge.

