

NO:	FORMULA OF COMPOUNDS-SET-1										
1	<p>The formulae of some ions are shown.</p> <table border="1" data-bbox="646 321 1110 602"> <thead> <tr> <th data-bbox="646 321 878 373">positive ion</th> <th data-bbox="878 321 1110 373">negative ion</th> </tr> </thead> <tbody> <tr> <td data-bbox="646 373 878 426">Ti<sup>4+</sup></td> <td data-bbox="878 373 1110 426">PO<sub>4</sub><sup>3-</sup></td> </tr> <tr> <td data-bbox="646 426 878 478">Al<sup>3+</sup></td> <td data-bbox="878 426 1110 478">SO<sub>4</sub><sup>2-</sup></td> </tr> <tr> <td data-bbox="646 478 878 531">Mg<sup>2+</sup></td> <td data-bbox="878 478 1110 531">NO<sub>3</sub><sup>-</sup></td> </tr> <tr> <td data-bbox="646 531 878 602">K<sup>+</sup></td> <td data-bbox="878 531 1110 602">Cl<sup>-</sup></td> </tr> </tbody> </table> <p>Which formula is <b>not</b> correct?</p> <p><b>A</b> Al<sub>3</sub>(SO<sub>4</sub>)<sub>2</sub>      <b>B</b> K<sub>3</sub>PO<sub>4</sub>      <b>C</b> Mg(NO<sub>3</sub>)<sub>2</sub>      <b>D</b> TiCl<sub>4</sub></p>	positive ion	negative ion	Ti <sup>4+</sup>	PO <sub>4</sub> <sup>3-</sup>	Al <sup>3+</sup>	SO <sub>4</sub> <sup>2-</sup>	Mg <sup>2+</sup>	NO <sub>3</sub> <sup>-</sup>	K <sup>+</sup>	Cl <sup>-</sup>
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Ms-1	A										
2	<p>Iron(III) chromate is a yellow solid. It contains the ions Fe<sup>3+</sup> and CrO<sub>4</sub><sup>2-</sup>.</p> <p>What is the formula of iron(III) chromate?</p> <p><b>A</b> FeCrO<sub>4</sub>      <b>B</b> Fe<sub>3</sub>(CrO<sub>4</sub>)<sub>2</sub>      <b>C</b> Fe<sub>2</sub>CrO<sub>4</sub>      <b>D</b> Fe<sub>2</sub>(CrO<sub>4</sub>)<sub>3</sub></p>										
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