
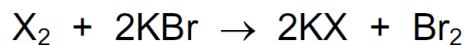


NO:	RELATIVE MOLECULAR MASS-RELATIVE FORMULA MASS-RELATIVE ATOMIC MASS-SET-2
1	<p>The diagram shows a model of a molecule of an organic acid.</p>  <p>What is the relative molecular mass of this acid?</p> <p><b>A</b> 11                      <b>B</b> 40                      <b>C</b> 58                      <b>D</b> 74</p>
Ms-1	D
2	<p>A compound has the formula <math>\text{CH}_3\text{CO}_2\text{H}</math>.</p> <p>How should the relative molecular mass, <math>M_r</math>, of this compound be calculated?</p> <p><b>A</b> <math>12 + 1 + 16</math></p> <p><b>B</b> <math>3(12 + 1) + 2(12 + 16) + 1</math></p> <p><b>C</b> <math>(4 \times 12) + (2 \times 1) + 16</math></p> <p><b>D</b> <math>(2 \times 12) + (4 \times 1) + (2 \times 16)</math></p>
Ms-2	D

3

Element X is in Group VII of the Periodic Table.

It reacts with aqueous potassium bromide as shown.



Which statements about X are correct?

	relative atomic mass	reactivity
<b>A</b>	greater than that of bromine	less reactive than bromine
<b>B</b>	greater than that of bromine	more reactive than bromine
<b>C</b>	less than that of bromine	less reactive than bromine
<b>D</b>	less than that of bromine	more reactive than bromine

Ms-3

D

4

Which relative molecular mass,  $M_r$ , is **not** correct for the molecule given?

	molecule	$M_r$
<b>A</b>	ammonia, $\text{NH}_3$	17
<b>B</b>	carbon dioxide, $\text{CO}_2$	44
<b>C</b>	methane, $\text{CH}_4$	16
<b>D</b>	oxygen, $\text{O}_2$	16

Ms-4

D

5	<p>Caesium chloride and rubidium bromide are halide compounds of Group I elements.</p> <p>Caesium chloride has the formula .....1....., a relative formula mass .....2..... that of rubidium bromide and bonds that are .....3..... .</p> <p>Which words correctly complete gaps 1, 2 and 3?</p> <table border="1" data-bbox="269 436 1052 705"> <thead> <tr> <th></th> <th>1</th> <th>2</th> <th>3</th> </tr> </thead> <tbody> <tr> <td><b>A</b></td> <td><math>\text{CaCl}</math></td> <td>different from</td> <td>ionic</td> </tr> <tr> <td><b>B</b></td> <td><math>\text{CaCl}</math></td> <td>the same as</td> <td>covalent</td> </tr> <tr> <td><b>C</b></td> <td><math>\text{CsCl}</math></td> <td>different from</td> <td>ionic</td> </tr> <tr> <td><b>D</b></td> <td><math>\text{CsCl}</math></td> <td>the same as</td> <td>covalent</td> </tr> </tbody> </table>		1	2	3	<b>A</b>	$\text{CaCl}$	different from	ionic	<b>B</b>	$\text{CaCl}$	the same as	covalent	<b>C</b>	$\text{CsCl}$	different from	ionic	<b>D</b>	$\text{CsCl}$	the same as	covalent
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Ms-5	C																				
6	<p>Iron forms an oxide with the formula <math>\text{Fe}_2\text{O}_3</math>.</p> <p>What is the relative formula mass of this compound?</p> <p><b>A</b> 76                      <b>B</b> 100                      <b>C</b> 136                      <b>D</b> 160</p>																				
Ms-6	D																				
7	<p>The relative atomic mass of chlorine is 35.5.</p> <p>When calculating relative atomic mass, which particle is the mass of a chlorine atom compared to?</p> <p><b>A</b> a neutron</p> <p><b>B</b> a proton</p> <p><b>C</b> an atom of carbon-12</p> <p><b>D</b> an atom of hydrogen-1</p>																				
Ms-7	C																				