

NO:	FINDING CONCENTRATION-SET-1
1	<p>25 cm³ of 0.1 mol/dm³ hydrochloric acid exactly neutralise 20 cm³ of aqueous sodium hydroxide.</p> <p>The equation for this reaction is:</p> $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ <p>What is the concentration of the sodium hydroxide solution?</p> <p>A 0.080 mol/dm³ B 0.800 mol/dm³ C 0.125 mol/dm³ D 1.25 mol/dm³</p>
Ms-1	C
2	<p>A solution of ethanoic acid, CH₃COOH, has a concentration of 2 mol/dm³.</p> <p>Which statement about this solution is correct?</p> <p>A 20 g of ethanoic acid is dissolved in 10 cm³ of water. B 30 g of ethanoic acid is dissolved in 250 cm³ of water. C 60 g of ethanoic acid is dissolved in 1 dm³ of water. D 120 g of ethanoic acid is dissolved in 2 dm³ of water.</p>
Ms-2	B
3	<p>4.00 g of solid sodium hydroxide is added to water to make a solution with a concentration of 0.200 mol/dm³.</p> <p>What is the volume of water used?</p> <p>A 0.5 cm³ B 20 cm³ C 500 cm³ D 2000 cm³</p>
Ms-3	C

4	<p>A tablet contains 0.080 g of ascorbic acid ($M_r = 176$).</p> <p>What is the concentration of ascorbic acid when one tablet is dissolved in 200 cm³ of water?</p> <p>A $9.1 \times 10^{-5} \text{ mol/dm}^3$</p> <p>B $4.5 \times 10^{-4} \text{ mol/dm}^3$</p> <p>C $9.1 \times 10^{-2} \text{ mol/dm}^3$</p> <p>D $2.3 \times 10^{-3} \text{ mol/dm}^3$</p>
Ms-4	D
5	<p>What is the concentration of a solution that contains 25.0 g NaOH in 500 cm³ of water?</p> <p>A 0.125 mol/dm³</p> <p>B 0.800 mol/dm³</p> <p>C 1.25 mol/dm³</p> <p>D 3.20 mol/dm³</p>
Ms-5	C