

NO:	FINDING VOLUME OF GAS PRODUCED-SET-1
1	<p>An excess of zinc is added to 100 cm³ of 1.0 mol/dm³ hydrochloric acid.</p> <p>The equation for the reaction is:</p> $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$ <p>What is the maximum volume of hydrogen evolved at room temperature and pressure?</p> <p>A 1.2 dm³ B 2.0 dm³ C 2.4 dm³ D 24 dm³</p>
2	<p>The equation for the reaction between potassium carbonate and nitric acid is shown.</p> $\text{K}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{KNO}_3 + \text{H}_2\text{O} + \text{CO}_2$ <p>Which volume of carbon dioxide is produced from 69 g of potassium carbonate?</p> <p>A 6 dm³ B 12 dm³ C 24 dm³ D 48 dm³</p>