# INERT ELECTRODES

**1** The results of experiments on electrolysis using inert electrodes are given in the table.

Complete the table; the first line has been completed as an example.

| electrolyte | change at negative electrode | change at positive electrode | change to electrolyte |
|---|---|---|---|
| molten lead(II) bromide | lead formed | bromine formed | used up |
| ..... ..... | potassium formed | iodine formed | used up |
| dilute aqueous sodium chloride | ..... | ..... | ..... ..... |
| aqueous copper(II) sulfate | ..... | ..... | ..... ..... |
| ..... ..... | hydrogen formed | bromine formed | potassium hydroxide formed |

[Total: 8]

**MARKING SCHEME:**

| | | |
|---|---|---|
| molten potassium iodide | **NOT** aqueous | [1] |
| hydrogen | | [1] |
| oxygen | | [1] |
| water used up **or** solution becomes more concentrated **or** sodium chloride remains<br>**NOT** no change | | [1] |
| If products are given as hydrogen, chlorine and sodium hydroxide then 2/3 | | |
| copper | | [1] |
| oxygen (and water) | | [1] |
| sulfuric acid | accept hydrogen sulfate | [1] |
| aqueous **or** dilute **or** concentrated potassium bromide<br>**accept** correct formulae | | [1] |

**[Total: 8]**