



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

**CHEMISTRY**

**0620/21**

Paper 2 Multiple Choice (Extended)

**May/June 2018**

**45 minutes**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)



**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

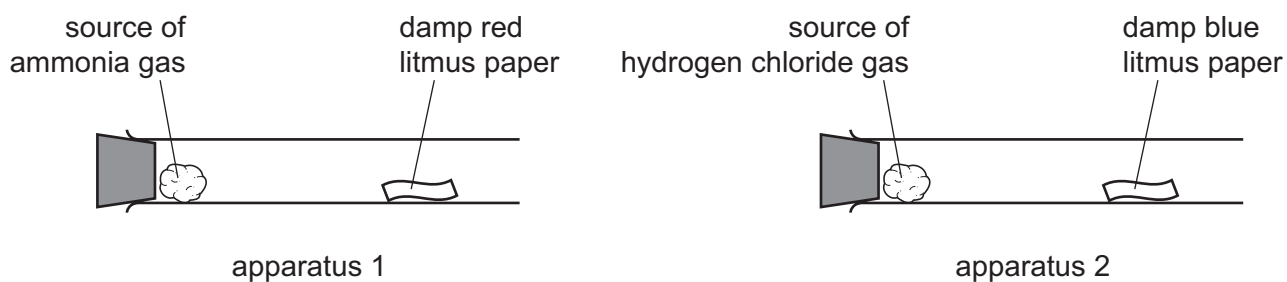
Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **13** printed pages and **3** blank pages.

- 1 A student investigated the diffusion of ammonia gas,  $\text{NH}_3$ , and hydrogen chloride gas,  $\text{HCl}$ .

Two sets of apparatus were set up as shown at room temperature and pressure.

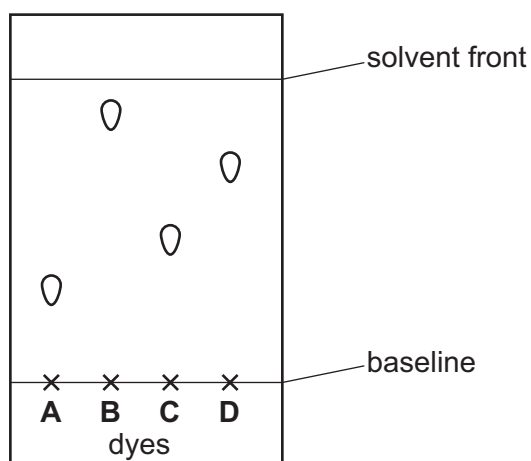


The damp red litmus paper in apparatus 1 changed colour after 30 seconds.

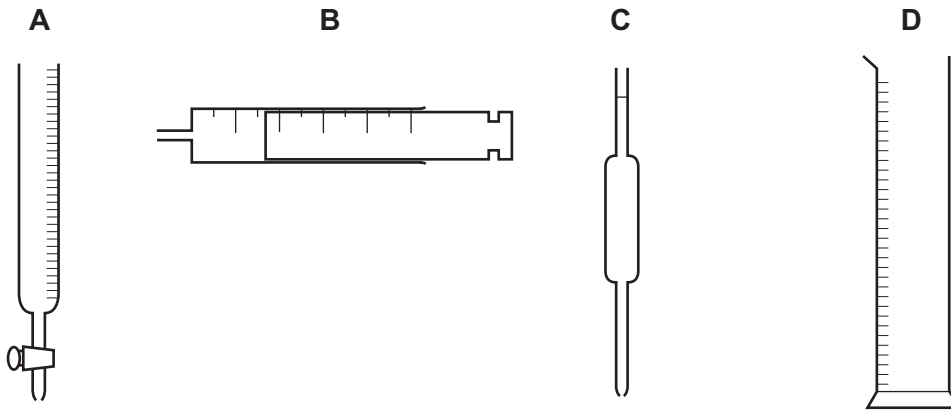
How long does it take for the damp blue litmus paper to change colour in apparatus 2?

- A** 64 seconds  
**B** 30 seconds  
**C** 21 seconds  
**D** The blue litmus paper would not change colour.
- 2 Chromatography is a technique used to separate coloured dyes.

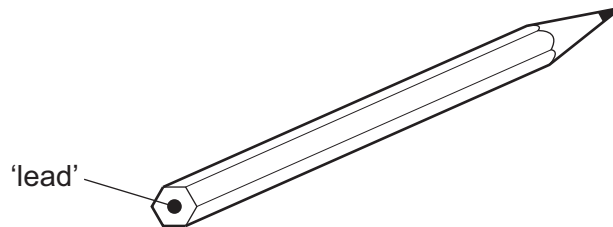
Which dye has an  $R_f$  value of 0.7?



3 Which piece of apparatus is used to measure exactly  $26.3 \text{ cm}^3$  of a liquid?



4 The 'lead' in a pencil is made of a mixture of graphite and clay.



When the percentage of graphite is increased, the pencil slides across the paper more easily.

Which statement explains this observation?

- A Graphite has a high melting point.
- B Graphite is a form of carbon.
- C Graphite is a lubricant.
- D Graphite is a non-metal.

- 5 Chlorine exists as two common isotopes,  $^{35}\text{Cl}$  and  $^{37}\text{Cl}$ .

Information about these two isotopes is shown.

|                  | number of protons | number of neutrons | number of electron shells |
|------------------|-------------------|--------------------|---------------------------|
| $^{35}\text{Cl}$ | 17                | 18                 | 3                         |
| $^{37}\text{Cl}$ | 17                | 20                 | 3                         |

Which statement explains why the two isotopes are of the same element?

- A** Both have the same number of electron shells.  
**B** Both have the same number of protons.  
**C** Both have 7 outer shell electrons.  
**D**  $^{37}\text{Cl}$  has 2 more neutrons than  $^{35}\text{Cl}$ .
- 6 Which substance is **not** a macromolecule?

- A** diamond  
**B** graphite  
**C** silicon(IV) oxide  
**D** sulfur

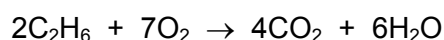
- 7 Copper is a metallic element.

Which statements about copper are correct?

- 1 Copper is malleable because layers of ions are in fixed positions and cannot move.
- 2 The structure of copper consists of negative ions in a lattice.
- 3 Copper conducts electricity because electrons can move through the metal.
- 4 Electrons hold copper ions together in a lattice by electrostatic attraction.

- A** 1 and 2      **B** 2, 3 and 4      **C** 2 and 3 only      **D** 3 and 4 only

- 8 The equation for the combustion of ethane is shown.



Which volume of carbon dioxide, at room temperature and pressure, is formed when 0.5 moles of ethane burn?

- A** 48 dm<sup>3</sup>      **B** 24 dm<sup>3</sup>      **C** 12 dm<sup>3</sup>      **D** 6 dm<sup>3</sup>

- 9 A solution of ethanoic acid,  $\text{CH}_3\text{COOH}$ , has a concentration of  $2 \text{ mol/dm}^3$ .

Which statement about this solution is correct?

- A 20 g of ethanoic acid is dissolved in  $10 \text{ cm}^3$  of water.
- B 30 g of ethanoic acid is dissolved in  $250 \text{ cm}^3$  of water.
- C 60 g of ethanoic acid is dissolved in  $1 \text{ dm}^3$  of water.
- D 120 g of ethanoic acid is dissolved in  $2 \text{ dm}^3$  of water.

- 10 Aqueous copper(II) sulfate is electrolysed using copper electrodes.

Which statement is correct?

- A A reduction reaction occurs at the positive electrode.
- B The blue colour of the solution becomes darker.
- C The concentration of copper ions in the solution decreases.
- D The mass of the negative electrode increases.

- 11 Dilute sulfuric acid is electrolysed using inert electrodes.

What are the ionic half-equations for the reactions that take place at each electrode?

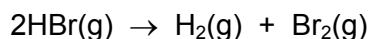
|   | positive electrode  | negative electrode  |
|---|---|---|
| A | $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$                        | $4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$ |
| B | $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$                        | $4\text{OH}^- + 4\text{H}^+ \rightarrow 4\text{H}_2\text{O}$              |
| C | $4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$ | $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$                        |
| D | $4\text{OH}^- + 4\text{H}^+ \rightarrow 4\text{H}_2\text{O}$              | $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$                        |

- 12 Plant cells use energy from sunlight for photosynthesis.

Which row describes and explains the energy change that occurs?

|   | type of energy change | explanation  |
|---|-----------------------|--|
| A | endothermic           | less energy is released making bonds than is absorbed to break bonds |
| B | endothermic           | more energy is released making bonds than is absorbed to break bonds |
| C | exothermic            | less energy is released making bonds than is absorbed to break bonds |
| D | exothermic            | more energy is released making bonds than is absorbed to break bonds |

- 13 Hydrogen bromide decomposes to form hydrogen and bromine. The equation is shown.



The bond energies are shown in the table. The reaction is endothermic.

| bond  | bond energy<br>in kJ/mol |
|-------|--------------------------|
| Br–Br | +193                     |
| H–Br  | +366                     |
| H–H   | +436                     |

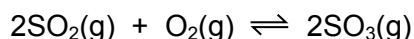
What is the energy change for the reaction?

- A** +263 kJ/mol    **B** +103 kJ/mol    **C** –103 kJ/mol    **D** –263 kJ/mol
- 14 Which row describes the effects of increasing both concentration and temperature on the collisions between reacting particles?

|          | increasing concentration   | increasing temperature   |
|----------|--|--|
| <b>A</b> | more collisions per second only  | more collisions per second only  |
| <b>B</b> | more collisions per second and more collisions with sufficient energy to react | more collisions per second only  |
| <b>C</b> | more collisions per second only  | more collisions per second and more collisions with sufficient energy to react |
| <b>D</b> | more collisions per second and more collisions with sufficient energy to react | more collisions per second and more collisions with sufficient energy to react |

- 15 The formation of sulfur trioxide is a reversible reaction.

The equation is shown.

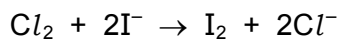


The forward reaction is exothermic.

Which conditions produce the highest equilibrium yield of sulfur trioxide?

|          | pressure | temperature |
|----------|----------|-------------|
| <b>A</b> | high     | high        |
| <b>B</b> | high     | low         |
| <b>C</b> | low      | high        |
| <b>D</b> | low      | low         |

16 Chlorine displaces iodide ions from potassium iodide.



What is the oxidising agent?

- A chloride ions
- B chlorine
- C iodide ions
- D iodine

17 Which statement about oxides is correct?

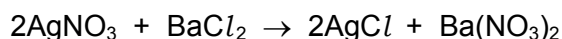
- A A solution of magnesium oxide has a pH less than pH 7.
- B A solution of sulfur dioxide has a pH greater than pH 7.
- C Magnesium oxide reacts with nitric acid to make a salt.
- D Sulfur dioxide reacts with hydrochloric acid to make a salt.

18 Which solution has the lowest pH?

- A 0.1 mol/dm<sup>3</sup> ammonia solution
- B 0.1 mol/dm<sup>3</sup> ethanoic acid
- C 0.1 mol/dm<sup>3</sup> lithium hydroxide
- D 0.1 mol/dm<sup>3</sup> nitric acid

19 A student mixes silver nitrate and barium chloride to form a white precipitate of silver chloride.

The equation is shown.



Which row describes the solubility of the salts?

|          | soluble   | insoluble   |
|----------|---|---|
| <b>A</b> | silver nitrate                                      | barium chloride, barium nitrate and silver chloride |
| <b>B</b> | silver nitrate and barium chloride                  | barium nitrate and silver chloride                  |
| <b>C</b> | silver nitrate, barium chloride and barium nitrate  | silver chloride                                     |
| <b>D</b> | silver nitrate, barium chloride and silver chloride | barium nitrate                                      |

20 Which methods are suitable for preparing **both** zinc sulfate and copper(II) sulfate?

- 1 reacting the metal oxide with warm dilute aqueous sulfuric acid
- 2 reacting the metal with dilute aqueous sulfuric acid
- 3 reacting the metal carbonate with dilute aqueous sulfuric acid

**A** 1, 2 and 3      **B** 1 and 2 only      **C** 1 and 3 only      **D** 2 and 3 only

21 Which element is in the same period of the Periodic Table as silicon?

- A** germanium  
**B** scandium  
**C** sodium  
**D** strontium

22 Which statement about the halogens is correct?

- A** A sample of bromine reacts with potassium chloride solution.  
**B** A sample of bromine reacts with potassium iodide solution.  
**C** A sample of chlorine has a higher density than a sample of bromine.  
**D** A sample of chlorine is a darker colour than a sample of bromine.

23 Which row shows the catalytic activity of transition elements and their compounds?

|          | catalytic activity of transition elements | catalytic activity of compounds of transition elements |
|----------|---|--|
| <b>A</b> | good                                      | good   |
| <b>B</b> | good                                      | poor   |
| <b>C</b> | poor                                      | good   |
| <b>D</b> | poor                                      | poor   |

24 The following statements are made about the metals copper, iron, magnesium and zinc.

- 1 Their oxides are acidic.
- 2 They all conduct electricity in the solid state.
- 3 They all have high melting points.
- 4 They all react with dilute acids to form hydrogen.

Which statements are correct?

**A** 1 and 2      **B** 1 and 4      **C** 2 and 3      **D** 3 and 4



25 Silver is a less reactive metal than cadmium.

Cadmium is a less reactive metal than barium.

Which statement is correct?

- A Barium does not react when heated with silver oxide.
- B Cadmium displaces barium from a solution of barium chloride.
- C Cadmium displaces silver from a solution of silver nitrate.
- D Cadmium reacts when heated with barium oxide.

26 Aluminium metal is extracted from aluminium oxide using electrolysis.

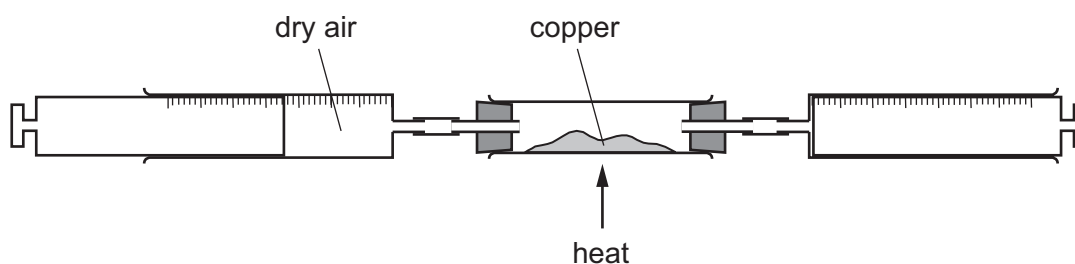
Which statement about the extraction process is **not** correct?

- A A large amount of electricity is required.
- B Molten cryolite is used to dissolve the aluminium oxide.
- C Oxygen gas is released which reacts to form carbon dioxide.
- D The negative electrodes burn away and have to be replaced.

27 Which statement explains why aluminium is used in the manufacture of aircraft?

- A It conducts heat well.
- B It has a low density.
- C It is a good conductor of electricity.
- D It is easy to recycle.

28 Dry air is passed over hot copper until all the oxygen has reacted.



The volume of gas at the end of the reaction is  $120\text{ cm}^3$ .

What is the starting volume of dry air?

- A  $132\text{ cm}^3$
- B  $152\text{ cm}^3$
- C  $180\text{ cm}^3$
- D  $570\text{ cm}^3$

29 A steel bicycle which had been left outdoors for several months was starting to rust.

What would **not** reduce the rate of corrosion?

- A Remove the rust and paint the bicycle.
- B Remove the rust and store the bicycle in a dry shed.
- C Remove the rust and wipe the bicycle with a clean, damp cloth.
- D Remove the rust and wipe the bicycle with an oily cloth.

30 Which statements about water are correct?

- 1 Household water contains dissolved salts.
- 2 Water for household use is filtered to remove soluble impurities.
- 3 Water is treated with chlorine to kill bacteria.
- 4 Water is used in industry for cooling.

- A 1, 2, 3 and 4
- B 1, 2 and 3 only
- C 1, 3 and 4 only
- D 2, 3 and 4 only

31 Ammonia is manufactured by reacting hydrogen with nitrogen in the Haber process.

Which row describes the sources of hydrogen and nitrogen and the conditions used in the manufacture of ammonia in the Haber process?

|          | source of hydrogen | source of nitrogen | temperature of reaction/°C | pressure of reaction/atm |
|----------|--------------------|--------------------|----------------------------|--------------------------|
| <b>A</b> | air                | natural gas        | 250                        | 2                        |
| <b>B</b> | air                | natural gas        | 250                        | 200                      |
| <b>C</b> | natural gas        | air                | 450                        | 2                        |
| <b>D</b> | natural gas        | air                | 450                        | 200                      |

32 Which statements about the carbon cycle are correct?

- 1 Carbon dioxide is added to the atmosphere by respiration.
- 2 Carbon dioxide is added to the atmosphere by combustion of coal.
- 3 Carbon dioxide is removed from the atmosphere by photosynthesis.

- A 1, 2 and 3
- B 1 and 2 only
- C 1 and 3 only
- D 2 and 3 only

- 33 Which statement about sulfur and its compounds is **not** correct?
- A Sulfur dioxide is used as a food preservative.
  - B Sulfur dioxide turns acidified aqueous potassium manganate(VII) from purple to colourless.
  - C Sulfur forms a basic oxide.
  - D Sulfur is used in the manufacture of sulfuric acid.
- 34 Which process is used to convert limestone (calcium carbonate) into lime?
- A electrolysis
  - B fractional distillation
  - C incomplete combustion
  - D thermal decomposition

- 35 What is **not** the correct use of the fraction named?

|   | name of fraction | use                    |
|---|------------------|------------------------|
| A | fuel oil         | making waxes           |
| B | gas oil          | fuel in diesel engines |
| C | kerosene         | jet fuel               |
| D | naphtha          | making chemicals       |

- 36 Which reaction is **not** a reaction which alkenes undergo?
- A bromination
  - B hydration
  - C hydrogenation
  - D hydrolysis
- 37 Which substances can be obtained by cracking hydrocarbons?
- A ethanol and ethene
  - B ethanol and hydrogen
  - C ethene and hydrogen
  - D ethene and poly(ethene)

38 Ethanol is produced by fermentation or from ethene.

What is a disadvantage of producing ethanol by fermentation?

- A Distillation is needed to purify the ethanol produced.
- B Fermentation uses glucose from plants.
- C Fermentation is catalysed by enzymes in yeast.
- D Fermentation occurs at a low temperature and pressure.

39 Which structural formula represents methyl propanoate?

- A  $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B  $\text{CH}_3\text{COOCH}_2\text{CH}_2\text{CH}_3$
- C  $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_3$
- D  $\text{HCOOCH}_2\text{CH}_2\text{CH}_3$

40 Which row describes addition polymerisation and condensation polymerisation?

|          | addition polymerisation   | condensation polymerisation   |
|----------|---|---|
| <b>A</b> | monomers have a C=C double bond and the polymer is the only product | monomers have a C=C double bond and the polymer is the only product |
| <b>B</b> | monomers have a C=C double bond and the polymer is the only product | the monomers react to form the polymer and a small molecule         |
| <b>C</b> | the monomers react to form the polymer and a small molecule         | monomers have a C=C double bond and the polymer is the only product |
| <b>D</b> | the monomers react to form the polymer and a small molecule         | the monomers react to form the polymer and a small molecule         |





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## The Periodic Table of Elements

|                            |                             | Group   |                             |                                 |                             |                              |                            |                           |                              |                                |                               |                               |                             |                               |                            |                           |                             |                           |                            |                           |                           |  |
|----------------------------|-----------------------------|---|-----------------------------|---------------------------------|-----------------------------|------------------------------|----------------------------|---------------------------|------------------------------|--------------------------------|-------------------------------|-------------------------------|-----------------------------|-------------------------------|----------------------------|---------------------------|-----------------------------|---------------------------|----------------------------|---------------------------|---------------------------|--|
| I                          | II                          | III   | IV                          | V                               | VI                          | VII                          | VIII                       |                           |                              |                                |                               |                               |                             |                               |                            |                           |                             |                           |                            |                           |                           |  |
| 3<br>Li<br>lithium<br>7    | 4<br>Be<br>beryllium<br>9   | 11<br>Na<br>sodium<br>23  | 12<br>Mg<br>magnesium<br>24 | 19<br>K<br>potassium<br>39      | 20<br>Ca<br>calcium<br>40   | 21<br>Sc<br>scandium<br>45   | 22<br>Ti<br>titanium<br>48 | 23<br>V<br>vanadium<br>51 | 24<br>Cr<br>chromium<br>52   | 25<br>Mn<br>manganese<br>55    | 26<br>Fe<br>iron<br>56        | 27<br>Co<br>cobalt<br>59      | 28<br>Ni<br>nickel<br>59    | 29<br>Cu<br>copper<br>64      | 30<br>Zn<br>zinc<br>65     | 31<br>Ga<br>gallium<br>70 | 32<br>Ge<br>germanium<br>73 | 33<br>As<br>arsenic<br>75 | 34<br>Se<br>selenium<br>79 | 35<br>Br<br>bromine<br>80 | 36<br>Kr<br>krypton<br>84 |  |
| 37<br>Rb<br>rubidium<br>85 | 38<br>Sr<br>strontium<br>88 | 39<br>Y<br>yttrium<br>89  | 57-71<br>lanthanoids        | 72<br>Hf<br>hafnium<br>178      | 73<br>Ta<br>tantalum<br>181 | 74<br>W<br>tungsten<br>184   | 75<br>Re<br>rhenium<br>186 | 76<br>Os<br>osmium<br>190 | 77<br>Ir<br>iridium<br>192   | 78<br>Pt<br>platinum<br>195    | 79<br>Au<br>gold<br>197       | 80<br>Hg<br>mercury<br>201    | 81<br>Tl<br>thallium<br>204 | 82<br>Pb<br>lead<br>207       | 83<br>Bi<br>bismuth<br>209 | 84<br>Po<br>polonium<br>— | 85<br>At<br>astatine<br>—   | 86<br>Rn<br>radon<br>—    |                            |                           |                           |  |
| 87<br>Fr<br>francium<br>—  | 88<br>Ra<br>radium<br>—     | 89<br>Ac<br>actinium<br>—   | 89-103<br>actinoids         | 104<br>Rf<br>rutherfordium<br>— | 105<br>Db<br>dubnium<br>—   | 106<br>Sg<br>seaborgium<br>— | 107<br>Bh<br>bohrium<br>—  | 108<br>Hs<br>hassium<br>— | 109<br>Mt<br>meitnerium<br>— | 110<br>Ds<br>darmstadtium<br>— | 111<br>Rg<br>roentgenium<br>— | 112<br>Cn<br>copernicium<br>— | 114<br>Fl<br>flerovium<br>— | 116<br>Lv<br>livermorium<br>— |                            |                           |                             |                           |                            |                           |                           |  |
|                            |                             | 1<br>H<br>hydrogen<br>1   |                             |                                 |                             |                              |                            |                           |                              |                                |                               |                               |                             | 2<br>He<br>helium<br>4        |                            |                           |                             |                           |                            |                           |                           |  |
|                            |                             | Key<br>atomic number<br>atomic symbol<br>name<br>relative atomic mass |                             |                                 |                             |                              |                            |                           |                              |                                |                               |                               |                             |                               |                            |                           |                             |                           |                            |                           |                           |  |

|             |                              |                            |                                 |                              |                             |                             |                             |                               |                            |                               |                              |                           |                               |                              |                              |
|-------------|------------------------------|----------------------------|---------------------------------|------------------------------|-----------------------------|-----------------------------|-----------------------------|-------------------------------|----------------------------|-------------------------------|------------------------------|---------------------------|-------------------------------|------------------------------|------------------------------|
| lanthanoids | 57<br>La<br>lanthanum<br>139 | 58<br>Ce<br>cerium<br>140  | 59<br>Pr<br>praseodymium<br>141 | 60<br>Nd<br>neodymium<br>144 | 61<br>Pm<br>promethium<br>— | 62<br>Sm<br>samarium<br>150 | 63<br>Eu<br>europium<br>152 | 64<br>Gd<br>gadolinium<br>157 | 65<br>Tb<br>terbium<br>159 | 66<br>Dy<br>dysprosium<br>163 | 67<br>Ho<br>holmium<br>165   | 68<br>Er<br>erbium<br>167 | 69<br>Tm<br>thulium<br>169    | 70<br>Yb<br>ytterbium<br>173 | 71<br>Lu<br>lutetium<br>175  |
| actinoids   | 89<br>Ac<br>actinium<br>—    | 90<br>Th<br>thorium<br>232 | 91<br>Pa<br>protactinium<br>231 | 92<br>U<br>uranium<br>238    | 93<br>Np<br>neptunium<br>—  | 94<br>Pu<br>plutonium<br>—  | 95<br>Am<br>americium<br>—  | 96<br>Cm<br>curium<br>—       | 97<br>Bk<br>berkelium<br>— | 98<br>Cf<br>californium<br>—  | 99<br>Es<br>einsteinium<br>— | 100<br>Fm<br>fermium<br>— | 101<br>Md<br>mendelevium<br>— | 102<br>No<br>nobelium<br>—   | 103<br>Lr<br>lawrencium<br>— |

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).