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0620/33

May/June 2018

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 20.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **17** printed pages and **3** blank pages.

1 The names of seven gases are given.

ammonia
ethene
helium
hydrogen
hydrogen chloride
methane
nitrogen

(a) Answer the following questions about these gases.
Each gas may be used once, more than once or not at all.
State which gas:

(i) forms an acidic solution in water

..... [1]

(ii) is monatomic

..... [1]

(iii) forms 78% of clean dry air

..... [1]

(iv) is the main constituent of natural gas

..... [1]

(v) has molecules which each contain 14 protons.

..... [1]

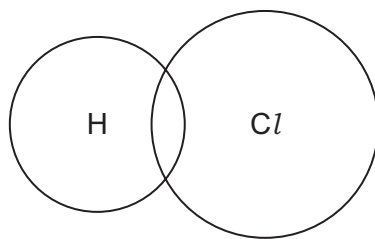
(b) (i) Give the names of **two** elements in the list.

..... and [1]

(ii) What is meant by the term *element*?

..... [1]

- (c) Complete the dot-and-cross diagram to show the electron arrangement in a molecule of hydrogen chloride. Show outer electrons only.



[2]

[Total: 9]

- 2 (a) The table shows the percentage by volume of each of the gases present in the exhaust gases from a petrol engine and a diesel engine.

name	percentage by volume from a petrol engine	percentage by volume from a diesel engine
nitrogen	72.00	67.00
carbon dioxide	14.00	
water vapour	12.00	11.00
carbon monoxide	1.50	0.05
oxides of nitrogen	0.25	0.70
hydrocarbons	0.24	0.22
sulfur dioxide	0.01	0.03
oxygen	0.00	9.00
	total 100.00	total 100.00

- (i) In the diesel engine, nitrogen, carbon dioxide and water vapour account for 90.00% of the exhaust gases by volume.

Calculate the percentage by volume of carbon dioxide in the exhaust gases from the diesel engine.

.....% [1]

- (ii) Describe **three** ways in which the composition of the exhaust gases from the petrol engine differ from the composition of the exhaust gases from the diesel engine.

1

2

3

[3]

- (iii) Give **one** adverse effect of sulfur dioxide on health.

..... [1]

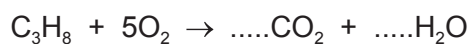
- (b) Complete the sentences about the separation of petroleum into fractions using words from the list.

boiling crystallisation density distillation kerosene poly(ethene)

Fractional of petroleum produces fractions such as gasoline, diesel oil and Within each fraction, the molecules have a particular range of points.

[3]

- (c) Balance the chemical equation for the complete combustion of propane.

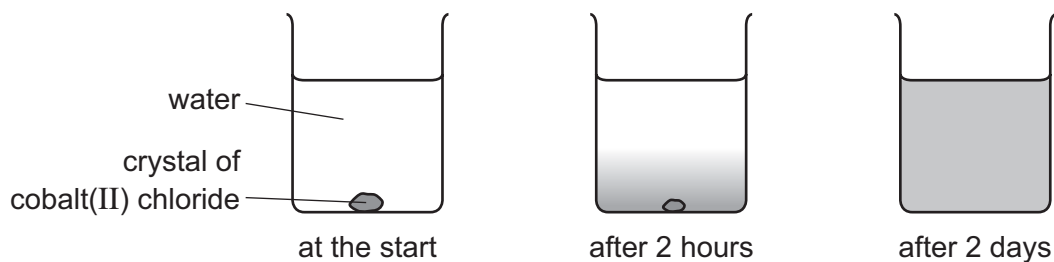


[2]

[Total: 10]

3 This question is about cobalt and its compounds.

- (a) A coloured crystal of cobalt(II) chloride was placed at the bottom of a beaker containing water. After 2 days, the colour had spread throughout the water.



Explain these observations using the kinetic particle model.

.....

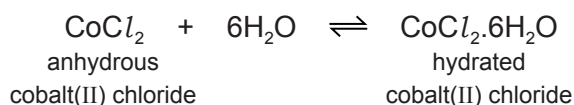
.....

.....

.....

..... [3]

- (b) Cobalt(II) chloride can be used to test for the presence of water.



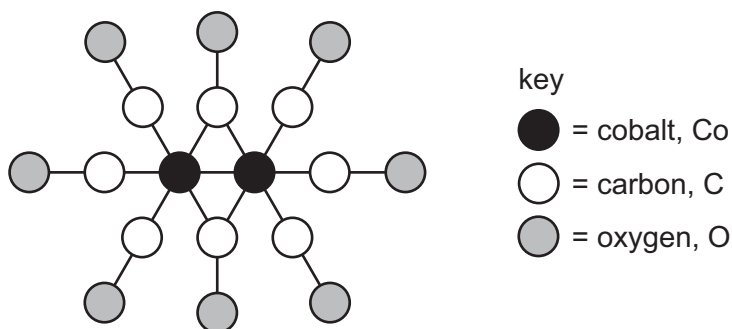
- (i) What is meant by the symbol $\rightleftharpoons$?

..... [1]

- (ii) Describe how the colour of anhydrous cobalt(II) chloride changes when water is added to it.

from to [2]

- (c) A compound of cobalt can be represented by the structure shown.



Deduce the molecular formula of this compound showing the number of cobalt, carbon and oxygen atoms.

..... [1]

- (d) The table compares the reactivity of cobalt with that of three other metals.

metal	reactivity with cold water	reactivity with steam
barium	reacts rapidly	
cobalt	no reaction	reacts slowly when heated
magnesium	reacts very slowly	reacts rapidly
zinc	no reaction	reacts easily when heated

Use this information to put the metals in order of their reactivity. Put the least reactive metal first.

least reactive $\xrightarrow{\hspace{10em}}$ most reactive

[2]

- (e) Cobalt is a transition element. Lithium is an element in Group I of the Periodic Table.

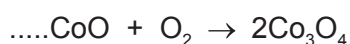
Describe **three** ways in which the properties of cobalt differ from those of lithium.

- 1
- 2
- 3

[2]

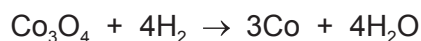
- (f) When cobalt(II) oxide, CoO, is heated in air, an oxide with the formula Co₃O₄ is formed.

Balance the chemical equation for this reaction.



[1]

- (g) When Co₃O₄ is heated with hydrogen, cobalt metal can be formed.



How does this equation show that Co₃O₄ is reduced?

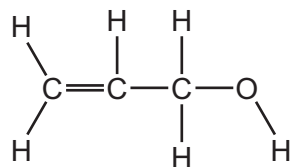
.....

..... [1]

[Total: 13]

4 This question is about alcohols.

(a) The structure of compound **P** is shown.



(i) What feature of the structure of compound **P** shows that it is unsaturated?

..... [1]

(ii) Describe how the colour of aqueous bromine changes when an excess of compound **P** is added to it.

from to [2]

(iii) The melting point of compound **P** is -129°C .
The boiling point of compound **P** is 97°C .

What is the physical state of compound **P** at -120°C ?
Explain your answer.

.....
..... [2]

(b) Ethanol is an alcohol with two carbon atoms.

(i) Draw the structure of ethanol showing all of the atoms and all of the bonds.

[1]

- (ii) Ethanol can be manufactured by fermentation.

Complete the sentences about fermentation using words from the list.

30 °C 300 °C distillation electrolysis
elements enzymes salt sugar

A flask containing solution and yeast is heated to

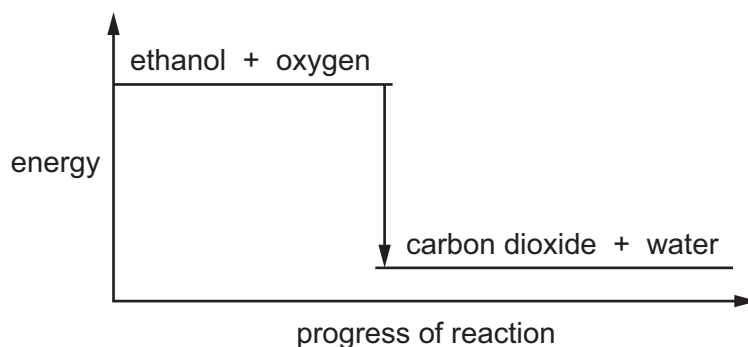
The yeast contains which catalyse the reaction. The ethanol is separated by [4]

- (iii) Ethanol can be used as a fuel.

Give **one** other use of ethanol in industry.

..... [1]

- (iv) The energy level diagram for the complete combustion of ethanol is shown.



Explain how this diagram shows that the reaction is exothermic.

..... [1]

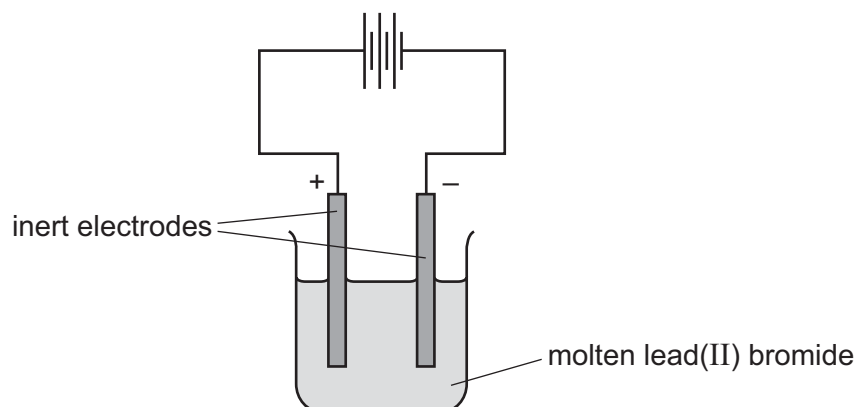
- (v) Carbon dioxide and water can be formed when ethanol burns in a limited supply of air.

State the name of **two** other products which can be formed when ethanol burns in a limited supply of air.

..... and [2]

[Total: 14]

- 5 (a) Molten lead(II) bromide can be electrolysed using the apparatus shown.



The negative electrode is called the cathode.

State the name of the positive electrode.

..... [1]

- (b) Predict the products of the electrolysis of molten lead(II) bromide at:

the negative electrode

the positive electrode.

[2]

- (c) Lead(II) bromide is insoluble in cold water.
Sodium bromide is soluble in cold water.

Describe how you could obtain crystals of sodium bromide from a mixture of solid lead(II) bromide and solid sodium bromide.

.....
.....
.....
..... [3]

- (d) When chlorine is bubbled through a colourless aqueous solution of sodium bromide, the solution turns orange-brown.

Name the orange-brown substance. Suggest, using ideas about reactivity of the halogens, why the reaction occurs.

orange-brown substance

why the reaction occurs

..... [2]

- (e) What is the colour of the precipitate formed when aqueous silver nitrate is added to aqueous sodium bromide?

..... [1]

- (f) Complete the sentence about the formation of a chloride ion from a chlorine atom.

A chloride ion is formed when a chlorine atom gains one [1]

[Total: 10]

6 This question is about isotopes.

(a) An atom of an isotope of oxygen is represented by the symbol shown.



Describe the structure of an atom of this isotope of oxygen.

In your answer, include:

- the position of the protons, neutrons and electrons in the atom
- the number of protons, neutrons and electrons present in the atom.

.....

.....

.....

.....

.....

.....

..... [5]

(b) Which **two** statements about isotopes are correct?

Tick **two** boxes.

Isotopes of the same element have different numbers of protons.

☐

Isotopes of the same element have different numbers of neutrons.

☐

Isotopes are always radioactive.

☐

The isotope ${}^{235}_{92}\text{U}$ is a source of energy.

☐

${}^{14}_6\text{C}$ and ${}^{14}_7\text{N}$ are isotopes of each other.

☐

[2]

(c) What is meant by the term *nucleon number*?

.....

..... [1]

[Total: 8]

7 This question is about Group I elements and their compounds.

(a) The properties of some Group I elements are shown in the table.

element	melting point / °C	relative hardness	observations when it reacts with cold water
lithium	181	23	
sodium	98	3	rapid bubbling but does not burst into flame
potassium	63		very rapid bubbling and bursts into flame
rubidium		1	bursts into flame and explodes

(i) Complete the table to estimate:

- the relative hardness of potassium
- the melting point of rubidium.

[2]

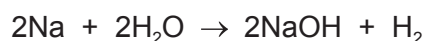
(ii) Use the information in the table to predict what you would observe when lithium reacts with cold water.

..... [1]

(b) Write the electronic structure of a sodium atom.

..... [1]

(c) When sodium reacts with water, an alkaline solution is formed.



(i) Use the information in the equation to explain why the solution formed is alkaline.

..... [1]

(ii) Describe how you could use a named indicator solution to show that the solution is alkaline.

.....
 [2]

(d) Sodium tetrahydridoborate, NaBH_4 , is an ionic compound.

Calculate the relative formula mass of NaBH_4 .

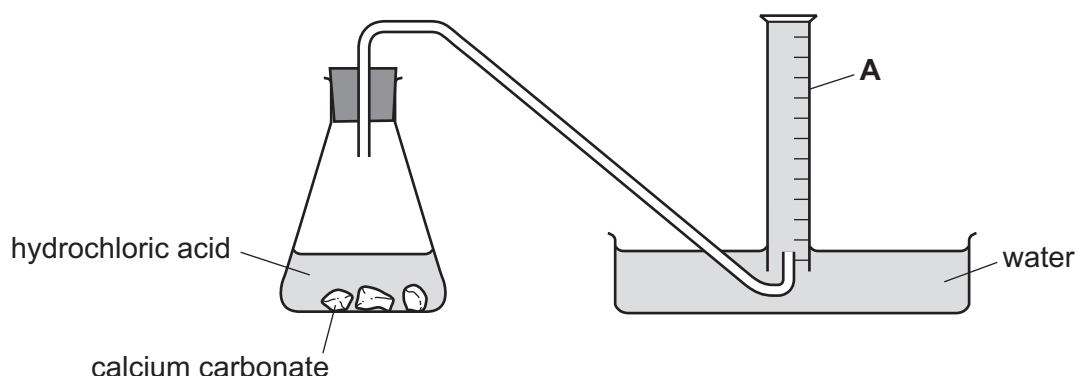
Show all your working.

Use your Periodic Table to help you.

relative formula mass = [2]

[Total: 9]

- 8 The apparatus shown is used to investigate the rate of reaction between calcium carbonate and hydrochloric acid at 30 °C.



- (a) Name the piece of apparatus labelled **A** in the diagram.

..... [1]

- (b) Describe how this apparatus can be used to find the rate of reaction.

.....

 [2]

- (c) The experiment is repeated at 20 °C. All other conditions are kept the same.

How does the decrease in temperature affect the rate of reaction?

..... [1]

- (d) The experiment is repeated using the same mass of different size pieces of calcium carbonate. All other conditions are kept the same.

The sizes of the pieces of calcium carbonate are:

- large pieces
- small pieces
- powder

Complete the table by writing the sizes of the pieces in the first column.

size of the pieces of calcium carbonate	initial rate of reaction in cm ³ gas/s
	5
	2
	12

[1]

- (e) Give the names of the **three** products formed when calcium carbonate reacts with hydrochloric acid.

1

2

3

[2]

[Total: 7]

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Group

Group																	
I	II											III	IV	V	VI	VII	VIII
1Hhydrogen1																	
Key																	
atomic number atomic symbol name relative atomic mass																	
3Li lithium 7	4Be beryllium 9																
11Na sodium 23	12Mg magnesium 24																
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131
55Cs caesium 133	56Ba barium 137	57–71lanthanoids	72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —
87Fr francium —	88Ra radium —	89–103actinoids	104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	114Fl flerovium —	116Lv livermorium —				
lanthanoids																	
57La lanthanum 139	58Ce cerium 140	59Pr praseodymium 141	60Nd neodymium 144	61Pm promethium —	62Sm samarium 150	63Eu europium 152	64Gd gadolinium 157	65Tb terbium 159	66Dy dysprosium 163	67Ho holmium 165	68Er erbium 167	69Tm thulium 169	70Yb ytterbium 173	71Lu lutetium 175			
actinoids																	
89Ac actinium —	90Th thorium 232	91Pa protactinium 231	92U uranium 238	93Np neptunium —	94Pu plutonium —	95Am americium —	96Cm curium —	97Bk berkelium —	98Cf californium —	99Es einsteinium —	100Fm fermium —	101Md mendelevium —	102No nobelium —	103Lr lawrencium —			