

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

COMBINED SCIENCE 0653/13

Paper 1 Multiple Choice October/November 2011

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

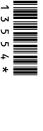
Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

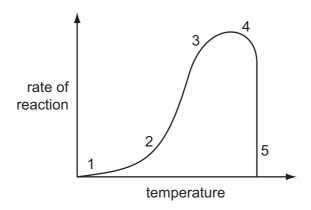


International Examinations

- 1 Which part of a plant cell is made of cellulose?
 - A cell membrane
 - B cell wall
 - **C** chloroplast
 - **D** nucleus
- 2 Which leaf tissue has specialised cells that surround stomata?
 - A epidermis
 - B palisade mesophyll
 - C phloem
 - **D** xylem
- 3 Which parts of a cell control its activities and control what enters and leaves it?

	controls cell's activities	controls what enters and leaves the cell
Α	chloroplast	cell surface membrane
В	chloroplast	cell wall
С	nucleus	cell surface membrane
D	nucleus	cell wall

4 The graph shows the effect of temperature on the rate of an enzyme-controlled reaction.



Where on the graph has all the enzyme been denatured?

- **A** 1
- **B** 2 and 3
- **C** 3 and 4
- **D** 5

5 Oxygenated blood returns to the heart from the lungs in vessel X and leaves the heart to circulate around the body in vessel Y.

What are X and Y?

	Х	Y
Α	aorta	pulmonary vein
В	pulmonary artery	vena cava
С	pulmonary vein	aorta
D	vena cava	pulmonary artery

6 When a leaf is photosynthesising, in which direction do gases diffuse through the stomata?

	carbon dioxide	oxygen
Α	in	in
В	in	out
С	out	in
D	out	out

7 What happens during digestion?

	large pieces of food are broken into small pieces	large molecules are broken into small molecules
Α	✓	✓
В	✓	x
С	X	✓
D	X	x

8 A species of animal reproduces both sexually and asexually.

Which offspring will be clones?

	offspring from sexual reproduction	offspring from asexual reproduction
Α	✓	✓
В	✓	x
С	×	✓
D	×	×

9 The table shows the level of alcohol in a person's blood after drinking two litres of beer.

time after drinking beer (hours)	alcohol in the blood (grams/dm³)
1	7
2	5
3	3
4	0

How long will it be (in hours) before the person's reaction time returns to normal?

- **A** 0 to 1
- **B** 1 to 2
- **C** 2 to 3
- **D** 3 to 4
- 10 Which method of family planning is also likely to reduce the risk of the spread of syphilis?
 - A condom
 - **B** intra-uterine device (IUD)
 - C pill
 - **D** sterilisation
- 11 Which process reduces soil erosion on hilly ground?
 - A cutting down the trees
 - **B** increasing the number of grazing animals
 - **C** ploughing up and down the hilly ground
 - **D** terracing the hilly ground
- **12** Albino humans cannot make any pigment in their skin.

A pale-skinned student, who is **not** an albino, sits in the sun on a number of days. The student's skin becomes suntanned (darker).

What causes this suntanning to happen?

- A the environment and the student's albino alleles
- **B** the environment and the student's non-albino alleles
- **C** the environment only
- D the student's genes only

13 The diagram shows a food chain.



Which types of energy are represented by the black arrows and by the white arrows?

	black arrows	white arrows
Α	chemical	heat
В	chemical	light
С	heat	chemical
D	light	chemical

14 Element X has a nucleon number of 40.

The electron arrangement of element X is 2,8,8.

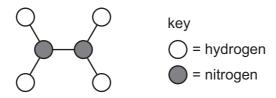
Which statements about element X are correct?

- 1 It has 40 neutrons in its nucleus.
- 2 It has 2 electrons in its outer shell.
- 3 It is unreactive.
- 4 It is in Group 0 of the Periodic Table.

A 1 and 2

- **B** 1 and 3
- **C** 2 and 4
- **D** 3 and 4

15 A model of a molecule is shown.

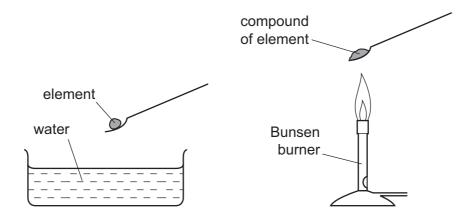


Which description and formula are correct for this molecule?

	description	formula
Α	compound	NH_2
В	compound	N_2H_4
С	mixture	NH_2
D	mixture	N_2H_4

16 In an experiment the elements calcium, copper, potassium and sodium were separately reacted with water.

In a second experiment a flame test was carried out on compounds of each of the elements.



Which row correctly shows the reaction of the elements with water and the colour of the flame?

	element	reaction with water	colour of the flame
Α	calcium	vigorous	green
В	copper	no reaction	red
С	potassium	vigorous	lilac
D	sodium	no reaction	yellow

- 17 Which two elements do **not** form an alloy?
 - A carbon and sulfur
 - B carbon and iron
 - C copper and zinc
 - **D** silver and gold
- **18** Sulfur dioxide is formed as a pollutant when fossil fuels are burned.

Which properties does sulfur dioxide have?

	toxic	acidic	corrosive
Α	✓	✓	✓
В	✓	✓	X
С	✓	X	X
D	X	X	X

19 Which equation is correctly balanced?

A
$$2Al + 3Cl_2 \rightarrow 2AlCl_3$$

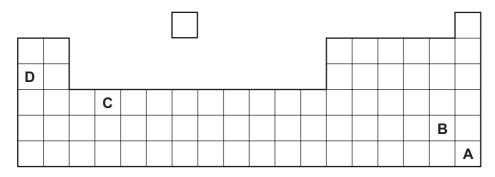
B Fe₂O₃ + 3C
$$\rightarrow$$
 2Fe + 3CO₂

C KC
$$l$$
 + Br₂ \rightarrow KBr + C l_2

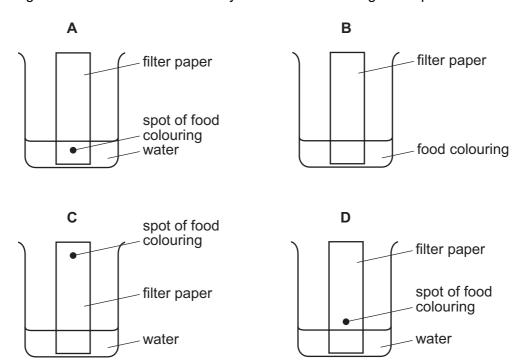
D Na +
$$H_2O \rightarrow NaOH + H_2$$

20 A soft metal reacts vigorously with cold water.

Which letter shows the position of this metal in the Periodic Table?



21 Which diagram shows how a mixture of dyes in a food colouring are separated?



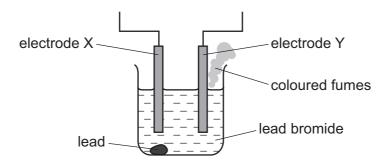
22 Ethene burns as shown.

$$C_2H_4(g) + 3O_2(g) \rightarrow 2CO_2(g) + 2H_2O(I)$$

What happens to ethene in this reaction?

- A decomposition
- **B** neutralisation
- **C** oxidation
- **D** reduction
- 23 The diagram shows the electrolysis of lead(II) bromide using inert electrodes.

Lead is formed at electrode X and coloured fumes at electrode Y.



Which statement about the electrolysis of lead(II) bromide is correct?

- A Electrode X is the anode.
- **B** The colour of the fumes is brown.
- **C** The lead(II) bromide is in aqueous solution.
- **D** The mass of the lead(II) bromide does not change during the reaction.
- **24** When compound X is added to pure water, the pH increases.

Which formula could **not** be a correct formula for X?

- A HNO₃
- **B** KOH
- **C** NaOH
- D NH₃

25 The structure of a molecule is shown.

Which term correctly describes this molecule?

- **A** hydrocarbon
- **B** monomer
- C petroleum
- **D** polymer

26 Many molecules of X combine to form a single molecule Y as shown in the equation.

$$n\: X\to Y$$

(n is a very large number)

Which terms best describe X and Y in this reaction?

	X	Υ
Α	fraction	monomer
В	monomer	fraction
С	monomer	polymer
D	polymer	fraction

- 27 Which change does **not** alter the rate of reaction between zinc and dilute sulfuric acid?
 - **A** addition of a catalyst
 - **B** change in concentration of the acid
 - C change in atmospheric pressure
 - D change in temperature
- **28** What is the meaning of the *weight* of an object?
 - A the density of the material from which it is made
 - **B** the force exerted on it by gravity
 - C the mass of the matter it contains
 - **D** the pressure it exerts on the ground

29 The table gives information about a liquid in a container.

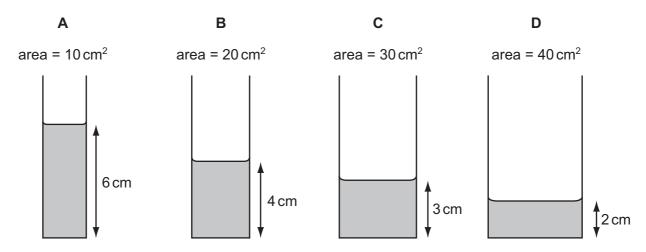
	Г
depth of liquid	10 cm
mass of liquid	30 g
temperature of liquid	25°C
volume of liquid	20 cm ³

What is the density of the liquid?

- **A** 0.33 cm/g
- **B** 1.2g/°C
- \mathbf{C} 1.5 g/cm³
- **D** 3.0 g/cm

30 Some water is poured into four tubes of different cross-sectional areas.

Which tube holds the largest volume of water?



31 An object travels 6.0 km in 2 minutes.

What is its speed?

- **A** 0.050 m/s
- **B** 3.0 m/s
- **C** 50 m/s
- **D** 3000 m/s

32 Which source releases energy by burning when it is used in the process of generating electricity?

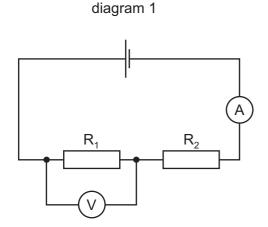
- A a fossil fuel
- **B** hydroelectric
- **C** nuclear
- **D** solar

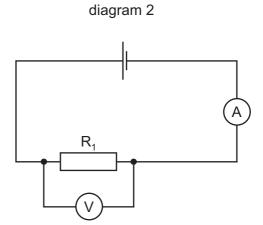
33 When flying, some birds use warm air currents to gain height.

What is the cause of these currents?

- **A** conduction
- **B** convection
- **C** evaporation
- **D** radiation
- **34** Why is a fuse used in an electric circuit in a house?
 - A to increase the resistance of the circuit
 - **B** to keep the power used to a minimum value
 - **C** to prevent a short circuit from occurring
 - **D** to stop the cables overheating
- 35 Diagram 1 shows two identical resistors R_1 and R_2 connected in series in a circuit.

 R_2 is then removed, as shown in diagram 2.





How do the readings on the ammeter and the voltmeter change when R₂ is removed?

	ammeter	voltmeter
Α	decreases	decreases
В	decreases increases	
С	increases	decreases
D	increases	increases

36 Which row shows two of the essential items used in the construction of a transformer?

	iron core	permanent magnet	primary coil	slip rings
Α	✓	✓		
В	✓		✓	
С		✓		✓
D			✓	✓

37 The diagram shows the spectrum of electromagnetic waves.

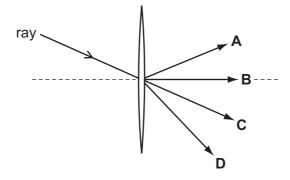
Which labelled region represents gamma rays?

A	micro waves	В	visible light	С	X-rays	D
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increasing frequency ----

- 38 Which is the best description of a wave that is a quiet, high-pitched sound?
 - A large amplitude and high frequency.
 - **B** large amplitude and low frequency.
 - **C** small amplitude and high frequency.
 - **D** small amplitude and low frequency.
- **39** A ray of light passes through the centre of a thin converging lens.

In which direction does the ray leave the lens?



40 Which nuclear process occurs in the Sun, and which process is used in a nuclear power station?

	in the Sun	in a nuclear power station
Α	fission	fission
В	fission	fusion
С	fusion	fission
D	fusion	fusion

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DATA SHEET
The Periodic Table of the Elements

	0	Helium	20 Neon 10 Neon 140 Ar Argon 18	84 Kry pton 36	Xe Xenon 54	Rn Radon 86		Lu Lutetium 71	
	₹		19 Fluorine 9 35.5 C1 C1	80 Br Bromine 35	127 I lodine 53	At Astatine 85		Yb Ytterbium	Š
	5		16 O Oxygen 8 32 Suffur 16	Selenium 34	128 Te Tellurium	Po Polonium 84		169 Tm Thullum	Md
	>		14 Nitrogen 7 31 Phosphorus 15	75 As Arsenic	122 Sb Antimony 51	209 Bi Bismuth 83		167 Er Erbium 68	Fm
	≥		12 Carbon 6 Silicon 14 Silicon 14	73 Ge Germanium 32	119 Sn ⊓in	207 Pb Lead		165 Ho Holmium 67	Es
	=		11 Boron 5 27 All Aluminium	70 Ga Gallium 31	115 I n Indium	204 T 1 Thallium 81		162 Dy Dysprosium 66	Č
				65 Zn Zinc 30	Cadmium 48	201 Hg Mercury 80		159 Tb Terbium 65	B
				64 Copper Copper 29	108 Ag Silver 47	197 Au Gold		157 Gd Gadolinium 64	CB
Group				59 Nickel 28	106 Pd Palladium 46	195 Pt Platinum 78		152 Eu Europium 63	Am
Ģ				59 Cobalt Cobalt	103 Rh Rhodium 45	192 I r Iridium 77		Sm Samarium 62	1
		1 Hydrogen		56 Fe Iron	Ru Ruthenium 44	190 Os Osmium 76		Pm Promethium 61	
				Manganese	Tc Technetium	186 Re Rhenium 75		144 Nd Neodymium 60	238 O
				52 Cr Chromium 24	96 Mo Molybdenum 42	184 W Tungsten 74		Pr Praseodymium 59	Pa
				51 Vanadium 23	93 Niobium 41	181 Ta Tantalum 73		140 Ce Cerium	²³²
				48 Ti Titanium 22	2 Zronium	178 Hf Hafnium 72			nic mass bol
				Scandium 21	89 ×	139 La Lanthanum 57 *	227 AC Actinium †	d series series	a = relative atomic massX = atomic symbol
	=		Beryllium 4 24 Mg Magnesium 12	40 Calcium 20	Strontium	137 Ba Barium 56	226 Ra Radium 88	*58-71 Lanthanoid series	e ×
	_		7 Lithium 3 23 Na Sodium 11	39 K Potassium 19	Rubidium	133 Caesium 55	Fr Francium 87	*58-71 L	Key

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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